MIPER

MOTHER THERESA INSTITUTE OF PHARMACEUTICAL EDUCATION & RESEARCH

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IV SEMESTER (II-B.PHARM)

MEDICINAL CHEMISTRY - I

PRACTICAL LAB MANUAL

1. ASSAY OF ASPIRIN

AIM:

To perform Assay of aspirin.

REFERENCE:

1. A textbook of Medicinal Chemistry-I,Pragi Arora,Varun Arora,Davinder Kumar,Page no:282,283.
2. Indian Pharmacopoeia Volume I 1996, Page No:69,70.
3. Pharmaceutical titrimetric analysis theory and practical,A.A Napoleon,Page No:11-17.

REQUIREMENTS:

Aspirin, Sodium hydroxide solution (0.5N),Hydrochloric acid (0.5N),Phenol red indicator, Burette, Conical flask, Funnel, Beaker etc.

PROCEDURE:

a) STANDARDIZATION OF 0.5M HYDROCHLORIC ACID

Weighed accurately 0.75g of Anhydrous sodium carbonate previously heated at 270^0C. Dissolve in 100 ml of water and added 0.1ml of methyl red solution. Added the titrant slowly from the burette with constant stirring until the solution becomes faintly pink. Heated the solution.Cool and continue.If pink colour fades on heating continue this process until a faint pink colour is no longer affected by continous boiling.

Each ml of 0.5M HCl = 0.026495g of Na_2CO_3

b) ASSAY OF ASPIRIN

Weighed accurately 1.5g of aspirin and dissolved in 15ml ethanol added 50ml of 0.5M sodium hydroxide boil gently for 10 minutes, cool and titrated the excess alkali with 0.5M HCl using phenol red solution as indicator. Perform a blank determination the difference between the titration represent the volume of sodium hydroxide consumed.

Each ml of 0.05M NaOH = 0.04504g of C_9H_8O

2.ASSAY OF PHENOBARBITONE

AIM :

To perform the Assay of Phenobarbitone.

REFERENCE:

1. A textbook of Medicinal Chemistry-I,Pragi Arora,Varun Arora,Davinder Kumar,Page no:282,283.
2. Indian Pharmacopoeia Volume III 2018,Page No:2899,2900.

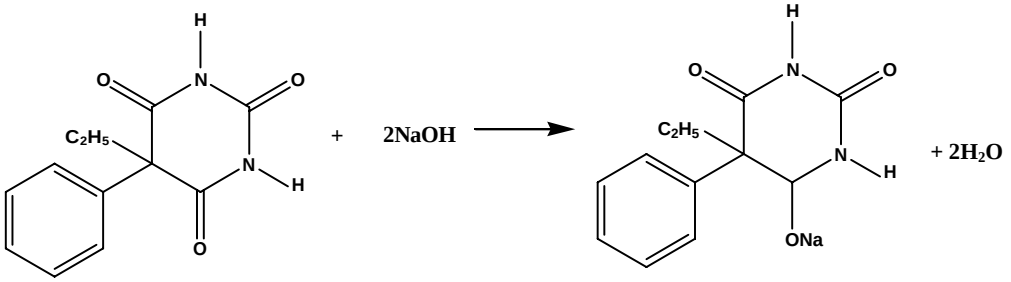
REQUIREMENTS:

Sodium hydroxide,aldehyde free ethanol,benzoic acid,thymolphthalein solution,

silver nitrate,pyridine and ether,conical flask,burette,beaker,pipette etc.

PRINCIPILE:

Phenobarbitone is assayed by non-aqueous titration.In this method,drug is dissolved in the pyridine and titrated with sodium hydroxide solution using thymolphthalein as an indicator.



Phenobarbitone **Sodium hydroxide** **Phenobarbitone sodium**

PROCEDURE:

a) STANDARDISATION OF SODIUM HYDROXIDE SOLUTION

Actually weighed 0.6g of benzoic acid and dissolved it in a mixture of 30ml of ethanol and 6ml of water and titrated with ethanolic sodium hydroxide solution using 0.2ml of thymolphthalein as indicator.

b) ASSAY OF PHENOBARBITONE

Weighed and powdered 20 tablets. Weighed a quantity of the powder containing about 0.1g (100 mg) of phenobarbitone in 5ml of pyridine add 0.25 ml of thymolphthalein solution and 10 ml of silver nitrate pyridine reagent and titrated with 0.1M ethanolic sodium hydroxide until a pure blue colour is obtained. Repeated the operation without the substance under examination. The difference between the titrations represents the amount of sodium hydroxide required.

Equivalent factor: 1ml of 0.1M ethanolic sodium hydroxide=0.01161g of $C_{12}H_{12}N_2O_3$

REPORT

The percentage purity of Phenobarbitone was found to be

CALCULATION

a) Standardization of 0.1M Sodium hydroxide solution

$$\textbf{Molarity of NaOH} = \frac{\textbf{Weight (W)}}{\textbf{Mol.wt of benzoic acid (122.12) X Volume (V)}}$$

Where,

W = Weight of benzoic acid (g)

V = Volume of NaOH solution consumed

b) Determination of Phenobarbitone

$$\% \text{ purity of phenobarbitone} = \frac{0.01161 \times V \times \text{Molarity(Calculated)} \times 100}{\text{Molarity (given)} \times W}$$

Where,

Molarity (calculated) = Molarity obtained from step (a)

V = Volume of Sodium hydroxide used

0.01161 is the equivalent factor

Molarity (given) = 0.1M

W = weight of sample

REPORT:

The percentage purity of Phenobarbitone was found to be =

3. ASSAY OF FUROSEMIDE

AIM:

To carry out the Assay of furosemide tablets.

REFERENCE:

1. A textbook of Medicinal Chemistry-I,Pragi Arora,Varun Arora,Davinder Kumar,Page no:288,289.
2. Indian Pharmacopoeia Volume III 2018,Page No:2899,2900.

REQUIREMENTS:

Furosemide, dimethyl formamide, sodium hydroxide, bromothymol blue indicator,0.1N oxalic acid, Phenolphthalein indicator, conical flask, burette, beaker, funnel etc.

PRINCIPLE

It is assayed by aqueous acid base titration between weak acid furosemide and strong alkali sodium hydroxide. In this assay protophilic solvent dimethyl formamide is used which enhances the acidity of furosemide so that it can be titrated with sodium hydroxide. To make the effect of acid impurities present negligible a solvent blank determination is carried out.

PREPARATION AND STANDARDIZATION OF STANDARD SOLUTIONS

a) SODIUM HYDROXIDE,XM

Solutions of any molarity xM may be prepared by dissolving 40x g of Sodium hydroxide in sufficient water to produce 1000ml.

b) STANDARDIZATION OF 0.1M SODIUM HYDROXIDE SOLUTION

Weighed accurately about 5g of potassium hydrogen phthalate previously dried at 120^0C for two hours dissolve in 75ml of carbon dioxide free water. Added 0.1ml of phenolphthalein solution and titrate with the sodium hydroxide until a permanent pink color is produced.

Each ml of 0.1M NaOH equivalent to 0.02042g of potassium hydrogen phthalate.

PROCEDURE:

a) ASSAY METHOD BY (NEUTRALIZATION TITRATION)

Weighed and powdered 20 tablets and Weighed accurately about a quantity of powder equivalent to 0.5g and dissolve in 40ml of dimethyl formamide and titrate with 0.1M sodium hydroxide using bromothymol blue as an indicator the end point shows the colour change from yellow to blue. Carry out a blank titration.

b) ASSAY METHOD BY (UV SPECTROPHOTOMETRY)

Weighed and powdered 20 tablets and Weigh accurately about a quantity of powder equivalent to 0.1g of furosemide and shake with 150ml of 0.1M sodium hydroxide for 10 minutes. Added sufficient 0.1M sodium hydroxide to produce 250ml and filter. Dilute 5ml to 200ml with 0.1M sodium hydroxide and measure the absorbance of the resulting solution at the maximum at about 271nm.Calculate the content of $C_{12}H_{11}ClN_2O_5S$ taking 580 as the value of A (1%, 1cm) at the maximum at about 271 nm.

REPORT:

The given sample contains.....mg of furosemide.

4. ASSAY OF IBUPROFEN

AIM:

To carry out the assay of Ibuprofen tablet

REFERENCE:

1. Indian Pharmacopoeia 2018 page no:2261-65.

REQUIREMENTS:

Ibuprofen,0.1N sodium hydroxide solution, phenolphthalein indicator, 0.1N oxalic acid solution, conical flask, burette, beaker etc.

PRINCIPLE:

Ibuprofen is determined by neutralization titration in which free carboxylic group is titrated with sodium hydroxide solution using phenolphthalein indicator. The amount of sodium hydroxide consumed in the reaction indicates the amount of ibuprofen present in the sample.

PROCEDURE:

PREPARATION AND STANDARDIZATION OF STANDARD SOLUTIONS

SODIUM HYDROXIDE.XM

Solutions of any molarity xM may be prepared by dissolving 40x g of Sodium hydroxide in sufficient water to produce 1000ml.

STANDARDIZATION OF 0.1M SODIUM HYDROXIDE SOLUTION

Weighed accurately about 5g of potassium hydrogen phthalate previously dried at 120^0C for two hours dissolve in 75ml of carbon dioxide free water. Added 0.1ml of phenolphthalein solution and titrate with the sodium hydroxide until a permanent pink color is produced.

Each ml of 0.1M NaOH equivalent to 0.02042g of potassium hydrogen phthalate.

Phenolphthalein solution

A 1.0%w/v solution of phenolphthalein in ethanol(95%).

PROCEDURE:

Weighed and powdered 20 tablets. Weighed a quantity of powder containing about 0.4g of ibuprofen, dissolve in 100ml of ethanol (95%) and titrated with 0.1M sodium hydroxide using 0.2ml of phenolphthalein solution as indicator. Perform a blank determination and make necessary correction.

Each ml of 0.1 M sodium hydroxide is equivalent to 0.02063 g of C_{13} H_{18} O_2

REPORT

The given sample contains.....mg of Ibuprofen.

5. ASSAY OF CHLORPROMAZINE

AIM: To carry out the Assay of Chlorpromazine.

REFERENCE: Indian Pharmacopoeia 2018 page no:1600-01.

REQUIREMENTS:

Perchloric acid (0.1M),Chlorpromazine, mercuric acetate solution(5% w/v in acetic acid),crystal violet solution (0.2%w/v in acetic acid),acetone,methyl orange indicator, conical flask, burette, beaker, potassium hydrogen phthalate, glacial acetic acid, crystal violet indicator.

PRINCIPLE:

Chlorpromazine is estimated by non-aqueous titration which is suitable for titration of weak acid and weak base. In this non aqueous solvent like perchloric acid is utilized as a titrant and methyl orange is used as an indicator. Mercuric acetate is added in the non-aqueous titration in order to remove the chloride ions. So as to prevent the interference of the chloride ion released by the titrant. The mercuric acetate replaces the halide ion in chlorpromazine with acetate ion which is a strong base. The end point is indicated by appearance of blue colour.

PROCEDURE:

a) STANDARDISATION OF PERCHLORIC ACID (0.1N)

Dissolved 0.5g of potassium hydrogen phthalate in 25ml of glacial acetic acid and added few drops of 5%w/v crystal violet indicator. Titrated the solution with 0.1N perchloric acid till blue green colour appears.

b) ASSAY OF CHLORPROMAZINE

Weighed accurately about 0.6g and dissolved in 200 ml of acetone. Added 15ml of mercuric acetate solution. Titrated with 0.1M perchloric acid, using a saturated solution of methyl orange in acetone as indicator. Perform a blank determination and make a necessary correction.

Each ml of 0.1M perchloric acid equivalent to 0.03553g of $C_{17}H_{19}ClN_2S,HCl$

REPORT:

The given sample contains.....mg of chlorpromazine.

6. ASSAY OF ATROPINE

AIM:

To carry out the assay of atropine.

REFERENCE:

1. Indian Pharmacopoeia 2018 page no:1600-01.

2. A textbook of Medicinal Chemistry-I,Pragi Arora,Varun Arora,Davinder Kumar,Page no:281,282

REQUIREMENTS:

Perchloric acid (0.1M),atropine,glacial acetic acid,crystal violet solution (0.2%w/v in acetic acid),acetone,methyl orange indicator, conical flask, burette, beaker.

PRINCIPLE:

Atropine is assayed by non-aqueous titration which is generally used for the titration of weak acid with weak base. In this titration non-aqueous solvent perchloric acid is used and crystal violet is used as an indicator. At the end point blue colour is obtained.

PROCEDURE:

a) STANDARDISATION OF PERCHLORIC ACID (0.1N)

Dissolved 0.5g of potassium hydrogen phthalate in 25ml of glacial acetic acid and few drops of 5%w/v crystal violet indicator. Titrated the solution with 0.1N perchloric acid till blue green colour appears.

b) ASSAY OF ATROPINE

Weighed accurately 400mg of atropine and dissolved it in 50ml of glacial acetic acid and added a drop of crystal violet indicator. Titrated this solution with 0.1N perchloric acid until green color is obtained end point.

7. DETERMINATION OF PARTITION COEFFICIENT OF BENZOIC ACID BETWEEN BENZENE AND WATER

AIM:

To determine partition coefficient of benzoic acid between benzene and water.

REFERENCE:

1. Medicinal chemistry – I, Mrs Sheethal V.Patil, Mrs.Swati G.Patil, Dr.Sunila T.Patil, Dr.Md.Rageeb, Md.Usman page no:270-272.
2. Textbook of Practical chemistry 2008,K.S mukherjee, Page no:293.

REQUIREMENTS:

Separating funnel(250ml),conical flask ,pipette, burette, stoppered bottle, Saturated solution of benzoic acid in benzene,benzene,0.01N NaOH, 0.1N NaOH and distilled water.

PRINCIPLE:

When a solute is shaken with two immiscible solvents it gets distributed between the solvents. This distribution of solute in two solvents depends on the solubility of the solute in two solvents. At the distribution equilibrium, the ratio of concentration of the solute in the two solvents is constant at a given temperature. The constant is called partition coefficient (K) or the distribution coefficient of the solute between the two solvents.

PROCEDURE:

Prepared the following mixtures in separating funnels:

Set I: 25ml water + 25ml of saturated solution of benzoic acid in benzene.

Set II: 25ml water + 20 ml saturated solution of benzoic acid in benzene + 5ml benzene.

Set III: 25ml water + 15ml saturated solution of benzoic acid in benzene + 10ml benzene.

Shaken the mixture in the separating funnel vigorously for about 30 minutes so that the benzoic acid gets distributed between the two solvents and the distribution equilibrium is reached. Allowed the flasks to stand for 10 minutes to separate into two clear layers (removed the stopper of the separating funnel and keep its mouth open during this period to facilitate the separation). Drain off the lower aqueous layers in 3 different stoppered dry bottles. (Discard the intermediate layer between the two phases).Benzene layer remains in the separating funnels. Using a dry pipette take 5ml of organic layer (Benzene) into a conical flask containing 10ml of water and titrate against 0.1N NaOH using Phenolphthalein as an indicator. The end point is indicated by the color change from colorless to pink. Pipette out

10ml of the aqueous layer using dry pipette and titrate it against NaOH solution using phenolphthalein as an indicator. End point is indicated by the color change from colorless to pink.

OBSERVATION

Set No.	V_{org}	V_{aq}	$N_{org} = C_{org}$	$N_{aq} = C_{aq}$	K	$\log C_{org}$	$\log C_{aq}$

Mean partition coefficient (K) =

Where,

V_{org} = Volume in ml of 0.1N Sodium hydroxide per 5ml of organic layer

V_{aq} = Volume in ml of 0.1N Sodium hydroxide per 5ml of aqueous layer

N_{org} = Normality of organic layer

N_{aq} = Normality of aqueous layer

C_{org} = Concentration of organic layer in g mole/lit

C_{aq} = Concentration of aqueous layer in g mole/lit

$K = C_{aq} / (C_{org})^{1/2}$ = Partition coefficient of benzoic acid in water and benzene

CALCULATIONS

Set I:

For organic layer

Normality of NaOH (N_1=0.1N)

Volume of Organic layer pipetted (V_2) = 5ml

N_1V_1 (Sodium hydroxide) = N_2V_2 (Organic layer)

$N_2 = \frac{0.1 \times V_1}{5} = N_{org}$

Similarly calculate concentration of benzoic acid in organic layer of sets II and III

For aqueous layer

Normality of NaOH (N_1=0.01N)

Volume of aqueous layer pipetted (V_2) = 5ml

N_1V_1 (Sodium hydroxide) = N_2V_2 (aqueous layer)

$N_2 = \frac{0.1 \times V_1}{5} = N_{aq}$

Similarly calculate concentration of benzoic acid in aqueous layer of sets II and III

Graph

Plot the graph of $\log C_{aq}$ Vs $\log C_{org}$

Partition coefficient (K) = $\frac{C_{aq}}{C_{org}^{1/2}}$

$\log C_{aq} = 1/n \log C_{org} + \log K$

Above equation is equation of a straight line (y = mx +c)

Result from graph

Slope (m) =1/n

Therefore n is nearly =

Substituting the value of slope of line in the equation

$\log C_{aq} = 1/n \log C_{org} + \log K$

$\log C_{aq} = \log C_{org} + \log K$

log K =

K =

REPORT:

1. Partition coefficient of benzoic acid between distilled water and benzene is…….. by calculation and .....by graph.

2. Since $C_{aq}/ C_{org}^{½}$ is practically constant benzoic acid exists as a dimer (n=2) in benzene.

3. Molecular condition of benzoic acid in benzene is 1/slope = n = ... ,molecules of benzoic acid associate in benzene.

8. DETERMINATION OF PARTITION COEFFICIENT OF BENZOIC ACID BETWEEN BENZENE AND WATER

AIM:

To determine partition coefficient of benzoic acid between benzene and water.

REFERENCE:

1. Medicinal chemistry – I, Mrs Sheethal V.Patil, Mrs.Swati G.Patil, Dr.Sunila T.Patil, Dr.Md.Rageeb, Md.Usman page no:270-272.
2. Textbook of Practical chemistry 2008,K.S mukherjee, Page no:293.

REQUIREMENTS:

Separating funnel(250ml),conical flask ,pipette, burette, stoppered bottle, Saturated solution of benzoic acid in benzene,benzene,0.01N NaOH, 0.1N NaOH and distilled water.

PRINCIPLE:

When a solute is shaken with two immiscible solvents it gets distributed between the solvents. This distribution of solute in two solvents depends on the solubility of the solute in two solvents. At the distribution equilibrium, the ratio of concentration of the solute in the two solvents is constant at a given temperature. The constant is called partition coefficient (K) or the distribution coefficient of the solute between the two solvents.

PROCEDURE:

Prepared the following mixtures in separating funnels:

Bottle no:	Volume of ether	Volume of Water	Volume of Benzoic acid
1	20ml	20ml	0.25g
2	20ml	20ml	0.5g

a) Standardization of 0.1M Sodium hydroxide

Dissolved 0.5g of potassium hydrogen phthalate in 75ml of water and added 0.1ml phenolphthalein as indicator and titrated using 0.1M NaOH.

Each ml of 0.1M Sodium hydroxide= 0.02042g of Potassium hydrogen phthalate.

b) Determination of Partition Coefficient

The bottles are well stoppered and shaken for 20minutes.After completion of shaking the mixture is allowed to stand 20-30 minutes. So that the two layer separate completely ether is lighter than water and exist at top while water remains at lower portion. Ether layer is evaporated to dryness and residue was dissolved in 10ml ethanol and titrates with 0.1M NaOH using phenolphthalein as indicator. Similarly 10ml of water is pipetted out and titrate with 0.1N NaOH.

Each ml of 0.1M NaOH = 0.01221g of benzoic acid.

REPORT :

The log P value of benzoic acid in bottle 1 =

The log P value of benzoic acid in bottle 2 =

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CALCULATIONS

Standarisation of Sodium hydroxide

Sl.No	Contents of conical flask	Burette reading		Volume of NaOH (ml)	Indicator	Endpoint
		Initial	Final			

Molarity of NaOH = $\frac{\text{Weight taken X Expected Molarity}}{\text{Titre value X Eq.Wt Factor}}$

Bottle 1(0.25g of benzoic acid)

Sl.No	Contents of Bottle 1	Burette reading		Volume of NaOH (ml)	Indicator	Endpoint
		Initial	Final			
1	Residue + 10ml ethanol (ether layer)					
2	10mlwater layer +10ml ethanol (water layer)					

Bottle 2(0.5g of benzoic acid)

Sl.No	Contents of Bottle 1	Burette reading		Volume of NaOH (ml)	Indicator	Endpoint
		Initial	Final			
1	Residue + 10ml ethanol (ether layer)					

2	10mlwater layer +10ml ethanol (water layer)					

Bottle 1(0.25g of benzoic acid)

Concentration of benzoic acid in aqueous layer =

$$\frac{\text{titre value X actual normality X Eq.wt factor}}{\text{Expected normality X Vol:pipetted}}$$

Concentration of benzoic acid in organic layer = $\frac{\text{titre value X actual normality X Eq.wt factor}}{\text{Expected normality X Vol:pipetted}}$

Bottle 2(0.5g of benzoic acid)

Concentration of benzoic acid in aqueous layer =

$$\frac{\text{titre value X actual normality X Eq.wt factor}}{\text{Expected normality X Vol:pipetted}}$$

Concentration of benzoic acid in organic layer = $\frac{\text{titre value X actual normality X Eq.wt factor}}{\text{Expected normality X Vol:pipetted}}$

Partition coefficient of benzoic acid in bottle 2 = $\frac{\text{titre value X actual normality X Eq.wt factor}}{\text{Expected normality X Vol:pipetted}}$

9. PREPARATION OF BENZIMIDAZOLE

AIM: To prepare and submit benzimidazole from o-phenylenediamine.

REFERENCE:

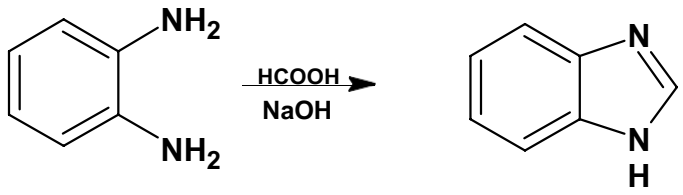
1. Medicinal chemistry – I, Mrs Sheethal V.Patil, Mrs.Swati G.Patil, Dr.Sunila T.Patil, Dr.Md.Rageeb, Md.Usman page no:233-235.

REQUIREMENTS:

Round bottom flask, Beaker, Measuring cylinder,Waterbath, Buchner funnel, O-phenylene diamine, Formic acid(90%),Sodium hydroxide (10%)

PRINCIPLE:

The preparation of benzimidazole can be done by reaction between O-phenylene diamine with formic acid in presence of base i.e sodium hydroxide.It is a condensation type of reaction in which o-phenylene diamine condensed with formic acid to give benzimidazole with removal of two molecules of water.



O-phenylene diamine Benzimidazole

PROCEDURE:

Placed 27g of O-phenylenediamine in a round bottomed flask of 250ml and added 17.5g (16ml) of 90% formic acid. Heated the mixture on a water bath at 100^0C for 2 hour. Cooled and added 10% sodium hydroxide solution slowly, with constant rotation of the flask, until the mixture is just alkaline to litmus. Filter off the synthesized crude benzimidazole by using the pump wash with ice cold water.

Recyrstallisation: Dissolved the synthesized product in 400ml of boiling water, added 2g of decolorizing carbon and digest for 15minutes. Filter rapidly through Buchner funnel and a flask at the pump. Cool the filtrate to about 10^0C, filter off the benzimidazole, wash with 25ml of cold water and dry at 100^0C. The yield of pure benzimidazole is 25g (85%), m.p 171-172^0C.

CALCULATION

Here limiting reagent is O-phenylene diamine; hence yield should be calculated from its amount taken.

Molecular formula of O-phenylene diamine = $C_6H_8N_2$

Molecular formula of benzimidazole = $C_7H_6N_2$

Molecular weight of O-phenylene diamine = 108g/mole

Molecular weight of benzilidazole = 118g/mole

108g of O-phenylene diamine forms 118g benzimidazole

Therefore, 27g O-phenylene diamine will form……(X) g benzimidazole

X = (118 X 27)/108 = 29.5g

Theoretical yield = 29.5g

Practical yield = ... g

% yield = $\frac{\text{(practical yield)}}{\text{(theoretical yield)}}$ X 100

REPORT:

Benzimidazole was synthesized from O-phenylene diamine and submitted.

10. SYNTHESIS OF BENZOTRIAZOLE

AIM:

To synthesize and submit benzotriazole from o-phenylene diamine and report its percentage yield.

REFERENCE:

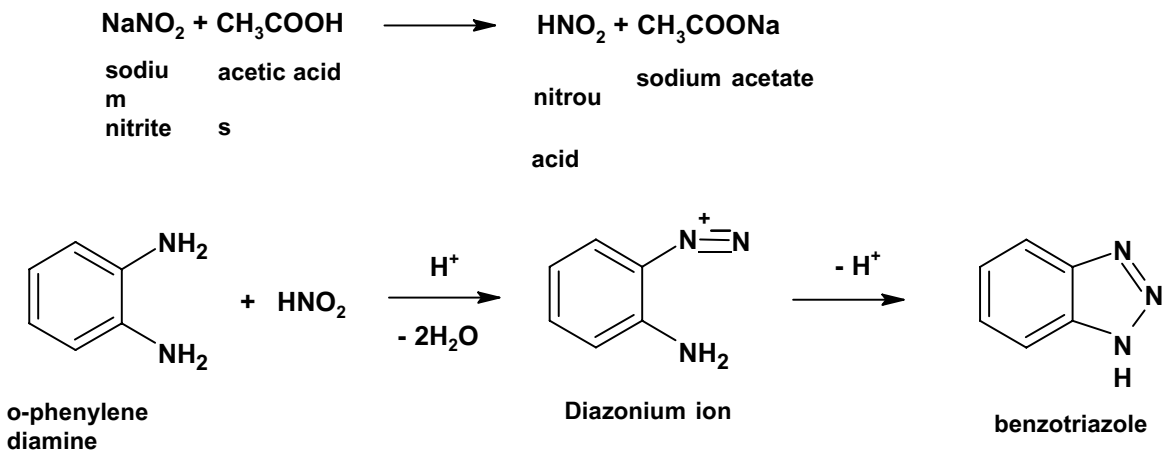
1. Practical medicinal chemistry by Dr. Devala Rao, page no:35.
2. Comprehensive practical organic chemistry by V.K.Ahluwalia and Renu Aggarval, page no:121.

CHEMICAL REQUIREMENTS:

o-phenylenediamine, glacial acetic acid, sodium nitrite

PRINCIPLE:

The sodium nitrite reacts with glacial acetic acid and liberates nitrous acid. The o-phenylene diamine reacts with nitrous acid and produce diazonium ion. When the structure and stereochemistry of diazonium ion are stable, intramolecular nitrogen coupling occurs and form benzotriazole directly.



CALCULATIONS

Molecular weight of o-phenylene diamine =

Molecular weight of benzotriazole =

---- g of o-phenylene diamine gives -----g of benzotriazole

1g of o-phenylene diamine =

=

---- g of o-phenylene diamine =

Theoretical yield =

Practical yield =

Percentage yield = $\frac{Practical\ yield}{theoretical\ yield}$ x 100

PROCEDURE:

Dissolve 1.3g of o-phenylenediamine in a mixture of 1.5ml of glacial acetic acid and 5ml water in a beaker. Stir until the solid dissolves, warm gently if necessary. Cool the solution to 15^0C. Stir well and add a solution of 2g of sodium nitrite in 2ml water. Reaction mixture become warm within 2-3 minutes and reaches a temperature of about 85^0C and then begins to cool. Colour changes from deep red to pale brown. Continue stirring for 15 minutes till the temperature fall about 35-40^0C. Thoroughly chill in ice bath for 30 minutes. Filter the product and wash with cold water.

USE:

Used in bulk drug industry as an important intermediate compound.

It is the basic nucleus present in anthelmintic drugs like mebendazole, thiabendazole etc.

REPORT:

Benzotriazole was prepared and submitted. The percentage yield was found to be ---------

11. SYNTHESIS OF 2,3-DIPHENYL QUINOXALINE

AIM:

To synthesize and submit 2,3-diphenyl quinoxaline from o-phenylenediamine and report its percentage yield.

REFERENCE:

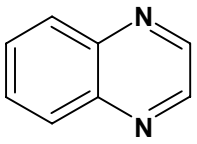
1) Vogels textbook of practical organic chemistry,5th edition, page no:90.
2) Comprehensive practical organic chemistry by V.K.Ahluwalia and Renu Aggarwal, page no:123.

CHEMICAL REQUIREMENTS:

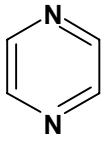
o-phenylenediamine, benzil, rectified spirit.

PRINCIPLE:

Quinoxalines are a type of heterocyclic compounds. They are also known as benzopyrazines.

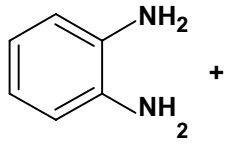


benzopyrazine



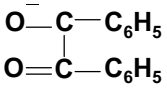
pyrazine

Generally quinoxaline is formed by the condensation of o-phenylenediamine with diketones. Here 2,3-diphenyl quinoxaline is prepared by treating o-phenylenediamine with benzil.

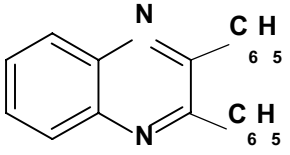


o-phenylene diamine

+



benzil



2,3,-diphenyl quinoxaline

CALCULATIONS

Molecular weight of 2, 3-diphenyl quinoxaline =

Molecular weight of o-phenylene diamine =

----- g of o-phenylene diamine gives ----- g of 2, 3-diphenyl quinoxaline

1g of o-phenylene diamine =

=

----- g of o-phenylene diamine =

Theoretical yield =

Practical yield =

Percentage yield = $\frac{Practical\ yield}{theoretical\ yield}$ x 100

=

PROCEDURE:

Add a solution of 1.1g of o-phenylenediamine in 8ml rectified spirit to a warm solution of 2.1g of benzil in 8ml rectified spirit. Warm the mixture for 30 minutes in a water bath. Add water dropwise until slight cloudiness persists. Cool the solution and filter the product.

USE:

Quinoxaline derivatives are used as antimicrobial agents like levomycin.

They are also used in dyes.

REPORT:

2,3-diphenyl quinoxaline was prepared and submitted.

The percentage yield was found to be ---------

12. SYNTHESIS OF PHENYTOIN

AIM:

To prepare and submit recrystallized dried product of phenytoin and calculate

(i) Percentage yield
(ii) Melting point

REFERENCE:

Medicinal chemistry theory and practical by and practical by Prof:K.Narayanan,Dr.Avjit Muzumder,Dr.L.K.Ghosh page no:9

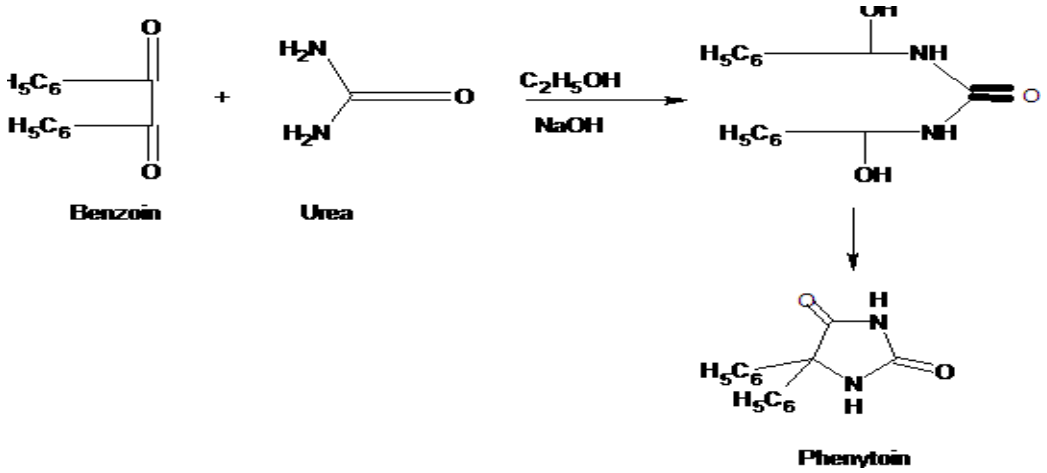
REQUIREMENTS:

Chemicals used**:** Urea, Nitric acid, Benzoin, Sodium hydroxide, ethanol, conc:HCl

Apparatus used**:** Round bottom flask ,reflex condenser, funnel, beaker, filter paper, glass rod.

PRINCIPLE:

Phenytoin is 5,5-diphenyl imidazoline 2,4-dione.Benzil react with urea in the presence of alkali and alcohol to give phenytoin by pinacolone rearrangement.



PROCEDURE:

a) Preparation of Benzil from Benzoin:

Place 2g of benzoin and 5ml of concentrated HNO_3 in a round bottom flask and heat on a boiling water bath till crystalline benzoin is replaced by oily benzil.Pour the mixture in to beaker of cold water with stirring the oily benzil crystallize in to yellow salt.

b) Preparation of phenytoin from benzil:

Place 1g benzil, 1g urea ,5ml 30% aqueous sodium hydroxide and 20ml ethanol in a round bottom flask which is attached to reflux condenser and boil for 2hours.Cool the mixture to attain room temperature.Pour the mixture to 100ml water and,mix and allow to stand for 15minutes.Filter to remove insoluble biproducts.Render the filtrate strongly acidic with concentrated HCl .Cool the filtrate in ice cold H_2O.Filter the precipitate product dry and submit.

IDENTIFICATION

Experiment	Observation	Inference
To the sample solution add hydrochloric acid	White precipitate	Presence of phenytoin
To the sample add pyridine and copper sulphate solution	Blue colour	Presence of phenytoin

REPORT:-

13. SYNTHESIS OF BENZOCAINE [ETHYL PARA AMINO BENZOATE]

AIM:

To synthesis recrystallized product of benzocaine from para amino benzoic acid and calculate

(i) Percentage yield
(ii) Melting point

REFERENCE:

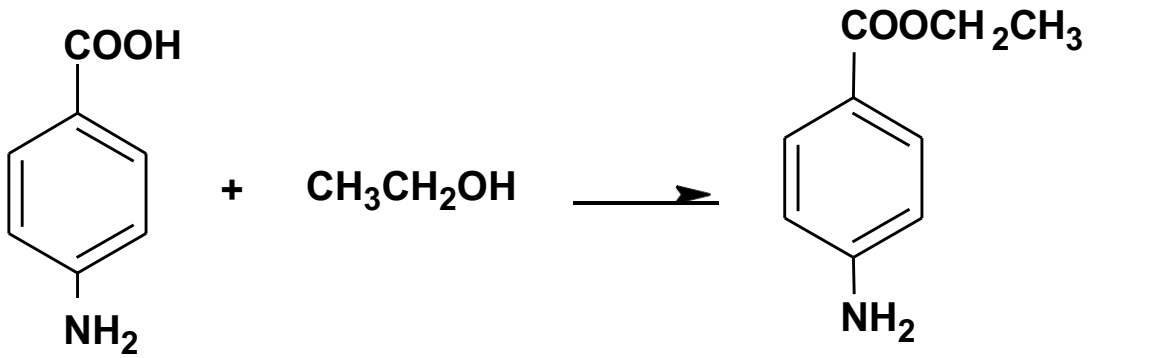
Medicinal chemistry theory and practical by and practical by Prof:K.Narayanan,Dr.Avjit Muzumder,Dr.L.K.Ghosh page no:22

REQUIREMENTS:

PABA, Conc:sulphuric acid, ethanol,reflux condenser, RB flask , beaker

PRINCIPLE:

Benzocaine is the ethyl ester of para amino benzoic acid(PABA).It can be prepaired from PABA and ethanol by fischer esterification.



PABA **Ethy para amino benzoate**

PROCEDURE:

To a 100ml RB flask, add 8ml of ethanol, 4.12g of para amino benzoic acid(PABA) and 1.2ml of conc:H_2SO_4 keep the mixture under reflux for 1hour up on cooling reaction mixture sets to a solid mass of hydrochloride of ethyl para amino benzoate. Pour the hot solution in to excess of water(no hydrochloride) add Na_2CO_3 to the clear solution until it is neutral to litmus. Filter wash and dry the product.

IDENTIFICATION

EXPERIMENT	OBSERVATION	INFERENCE
To the sample solution add sodium nitrite and con.HCl and cool the mixture. To this add a solution of beta naphthol in sodium hydroxide. Maintain the temperature at 0 to 5°	Deep red colour	Benzocaine confirmed

REPORT:-